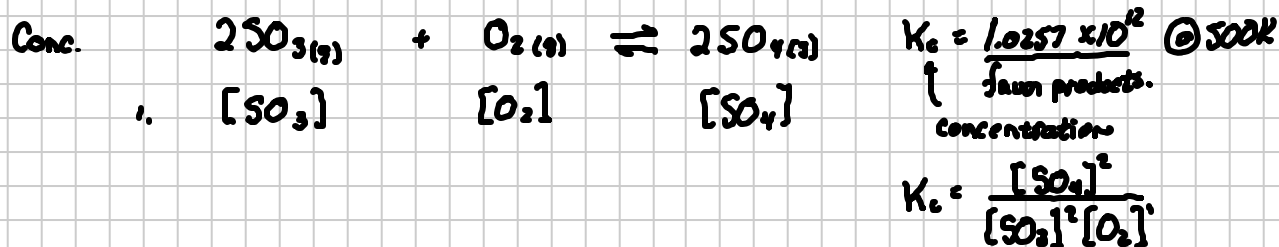


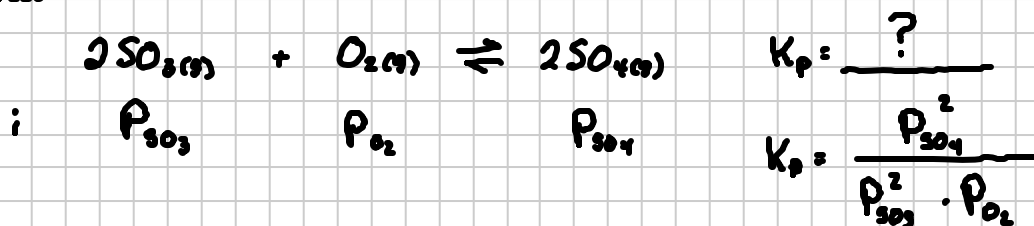
Lecture 8.4 K_c & K_p manipulations

Note Title

9/19/2011



Partial Pressure

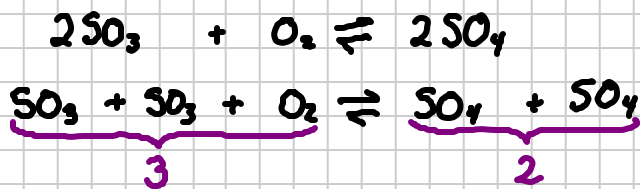


$$K_c \neq K_p$$

Conversion of K_p & K_c

$$K_c (RT)^{\Delta n} = K_p$$

Temp Kelvin.
 $\Delta n = ?$
 $0.0820578 \frac{\text{Latm}}{\text{molK}}$



$$\Delta n = \# \text{prod species} - \# \text{reactant species}$$

$$= 2 - 3$$

$$= -1$$

$$K_c (RT)^{\Delta n} = K_p$$

$$(1.0257 \times 10^{12}) (0.0820578 \cdot 500K)^{-1} = K_p$$

$$K_p = 2.50 \times 10^{10} \text{ Large... Product.}$$