

Empirical Formula of a Compound Pre-lab exercise: KEY

Show the calculations for the following questions:

1. What are the cautionary measures that you should take in handling the crucible in today's experiment?

FOLLOW ALL FIVE MEASURES DESCRIBED ON PAGE 2 OF PROTOCOL

2. What type of balance will you be using today, for mass measurements?

ANALYTICAL

3. What cautionary measures do you take in handling the balance?

FOLLOW ALL MEASURES DESCRIBED ON PAGE 2 OF PROTOCOL

4. In today's experiment, when magnesium burns in air, in addition to the oxide being formed, what other product will be formed?

MAGENSIUM NITRIDE (Mg_3N_2)

5. How do you convert this other compound to oxide?

ADD WATER

6. A 3.70 g sample of sodium is allowed to react completely with sulfur to form a sulfide which weighs 6.30 g. Calculate the following:

a) mass of sulfur ___**2.60 g S**___

b) moles of sulfur ___**0.0811 moles S**___

c) moles of sodium ___**0.161 moles Na**___